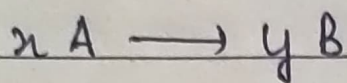


Question- In the following reaction



$$\log_{10} \left[\frac{-d[A]}{dt} \right] = \log_{10} \left[\frac{d[B]}{dt} \right] + 0.3010$$

'A' and 'B' respectively can be:

<1> n-butane and Iso butane

<2> C_2H_2 and C_2H_6

<3> C_2H_4 and C_4H_8

<4> N_2O_4 and NO_2

Answer- <3>

(JEE MAINS 2019, 12

April Morning)