

Exemplar problem 6

The magnitude of the surface tension of liquid depends on the attractive forces between the molecules. Arrange the following in increasing order of surface tension: Water, alcohol (C₂H₅OH) and hexane [CH₃(CH₂)₄CH₃].

Solution:

H-bonding is stronger in water than alcohol, so water has strong intermolecular attraction than alcohol. Increasing order of surface tension is – Hexane < alcohol < water.