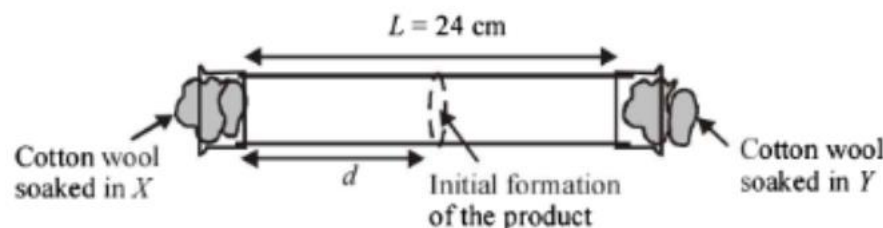


## Question 1

$X$  and  $Y$  are two volatile liquids with molar weights of  $10 \text{ g mol}^{-1}$  and  $40 \text{ g mol}^{-1}$  respectively. Two cotton plugs, one soaked in  $X$  and the other soaked in  $Y$ , are simultaneously placed at the ends of a tube of length  $L = 24 \text{ cm}$ , as shown in the figure. The tube is filled with an inert gas at 1 atmosphere pressure and a temperature of  $300 \text{ K}$ . Vapours of  $X$  and  $Y$  react to form a product which is first

observed at a distance  $d \text{ cm}$  from the plug soaked in  $X$ . Take  $X$  and  $Y$  to have equal molecular diameters and assume ideal behaviour for the inert gas and the two vapours.



- The value of  $d$  in cm (shown in the figure), as estimated from Graham's law, is *(JEE Adv. 2014)*
  - 8
  - 12
  - 16
  - 20
- The experimental value of  $d$  is found to be smaller than the estimate obtained using Graham's law. This is due to *(JEE Adv. 2014)*
  - Larger mean free path for  $X$  as compared to that of  $Y$
  - Larger mean free path for  $Y$  as compared to that of  $X$
  - Increased collision frequency of  $Y$  with the inert gas as compared to that of  $X$  with the inert gas
  - Increased collision frequency of  $X$  with the inert gas as compared to that of  $Y$  with the inert gas

- (c) According to Graham's law of diffusion, if all conditions are identical,

$$r = \frac{1}{\sqrt{M}}$$

As in this question, all conditions are identical for  $X$  and  $Y$ , then

$$\frac{r_x}{r_y} = \sqrt{\frac{M_y}{M_x}}$$

## Question 1

$$\frac{d}{24-d} = \sqrt{\frac{40}{10}} = 2$$

$$\Rightarrow d = 48 - 2d \Rightarrow 3d = 48 \Rightarrow d = 16 \text{ cm}$$

(d) The general formula of mean free path ( $\lambda$ ) is

$$\lambda = \frac{RT}{\sqrt{2}\pi d^2 N_A p}$$

(d = diameter of molecule, p = pressure inside the vessel)

Since d and p are same for both gases, ideally their  $\lambda$  are same. Hence it must be the higher drift speed of X due to which it is facing more collisions per second with the inert gas in comparison to gas Y. Hence X faces more resistance from inert gas than Y and hence covers lesser distance than that predicted by Graham's law.