

Question 2

One of the assumptions of the kinetic theory of gases is that there is no force of attraction between the molecules of a gas. State and explain the evidence that shows that the assumption is not applicable for real gases.

Solution:

Under low pressure and at high temperature the assumption made by kinetic theory is true. At high temperature, the molecules will be very far from each other and at low pressure, the volume of molecule become negligible, so they don't interact and hence the assumption is valid at high temperature.