**Question 5**

**The magnitude of the surface tension of liquid depends on the attractive forces between the molecules. Arrange the following in increasing order of surface tension: Water, alcohol ($C_2H_5OH$) and hexane [$CH_3(CH_2)_4CH_3$)].**

**Answer:**

H-bonding is stronger in water than alcohol, so water has strong intermolecular attraction than alcohol. Increasing order of surface tension is – Hexane< alcohol< water.