

Question 5

5. Rank the following in the increasing strength of intermolecular forces $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$, $\text{CH}_3\text{CH}_2\text{OH}$, He

Answer:

The only intermolecular forces present in helium gas are **dispersion forces**, which are the weakest.

The type of bonding present in n-butane is **London Dispersion Forces**. This is because this alkane is considered to be non-polar in nature and the London dispersion forces are considered to be weak forces.

$\text{CH}_3\text{CH}_2\text{OH}$ have both dispersion forces and hydrogen bonds

