

## Question 2

2. In which mixture do you expect to find ion–dipole forces between solute and solvent to exist:  $\text{CH}_3\text{OH}$  in water or  $\text{Ca}(\text{NO}_3)_2$  in water?

$\text{CH}_3\text{OH}$

- Covalent compound

$\text{Ca}(\text{NO}_3)_2$ :

- **Ionic compound**
- **Nitrates** ( $\text{NO}_3^-$ ) are soluble.
- This means that this compound forms ions

**$\text{Ca}(\text{NO}_3)_2$  in water has ion–dipole forces between solute and solvent**