

**JEE PREVIOUS YEARS QUESTION-1**

**1. ACCORDING TO THE KINETIC THEORY OF GASES, IN AN IDEAL GAS, BETWEEN TWO SUCCESSIVE COLLISIONS A GAS MOLECULE TRAVELS –**

**[AIEEE-2003]**

- A. In a straight-line path**
- B. With an accelerated velocity**
- C. In a circular path**
- D. In a wavy path**

**Solution:**

According to kinetic theory the gas molecules are in a state of constant rapid motion in all possible directions colliding in a random manner with one another and with the walls of the container and between two successive collisions molecules travel in a straight line path but show haphazard motion due to collisions.