

To check if a hydride is electron precise or not,

We need to check if the central atom of the molecule in its lewis structure satisfies octet rule or not.

Case 1: It satisfies octet rule: Hydride is electron precise.

Case 2: Electrons on the central element of the molecule are less than that required to satisfy octet rule: Hydride is electron deficient.

Case 3: Electrons on the central element of the molecule are more than that required to satisfy octet rule: Hydride is electron rich.