

On the basis of kinetic theory of gases, the gas exerts pressure because its molecules :

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- A continuously lose their energy till it reaches wall.
- B are attracted by the walls of container.
- C suffer change in momentum when impinge on the walls of container.
- D continuously stick to the walls of container.

On basis of $KT \propto$, gas exerts pressure because its molecules contain uniform speed, random motion and elastic collision with each other and also with the walls of container. As a result, they suffer momentum change on striking the container walls. \rightarrow [Option (c)]