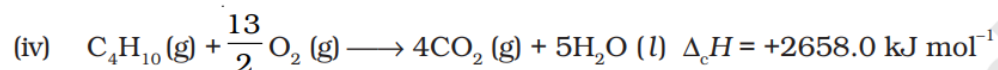
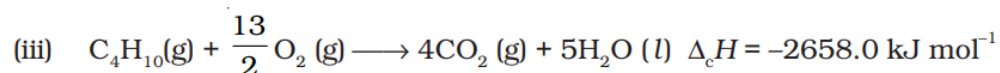
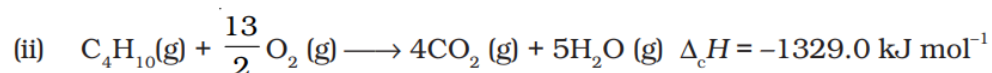
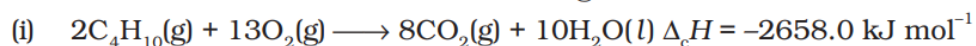


5. During complete combustion of one mole of butane, 2658 kJ of heat is released. The thermochemical reaction for above change is



Solution:

(iii) is the correct answer

Explanation:

As heat is released, hence, heat of combustion will be negative. By convention, heat of combustion is defined to the heat released for the complete combustion of a compound in its standard state to form stable products in their standard state: so water has to be in liquid state