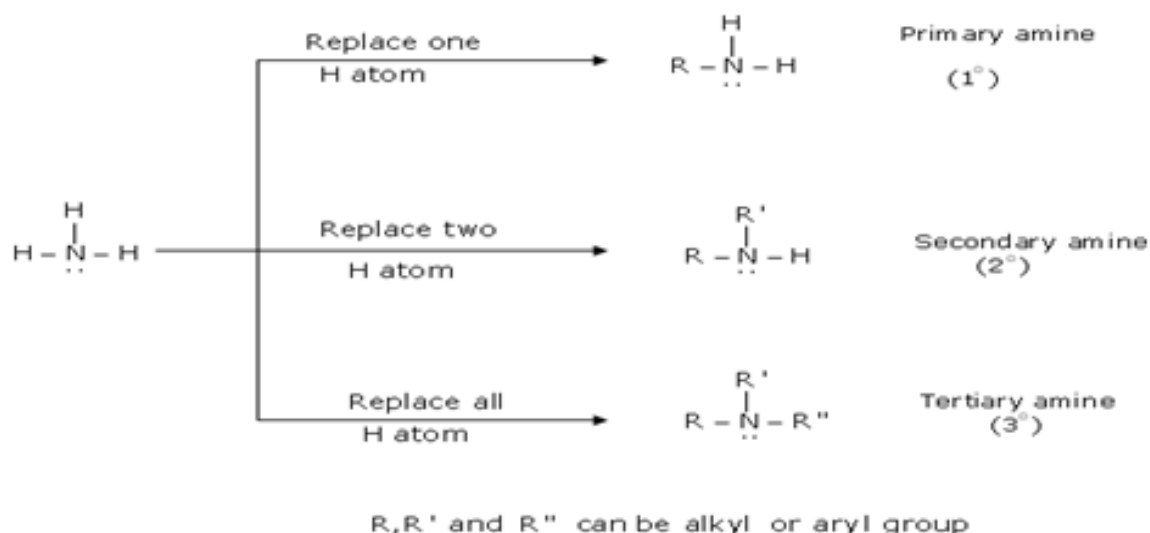


**NOTES**  
**CLASS XII CHAPTER-13**  
**AMINES**

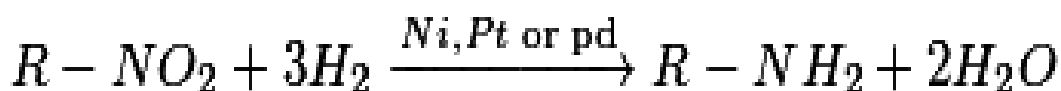
**1. Amines** are the derivatives of ammonia in which one or more hydrogen atoms have been replaced by alkyl groups.

**2. Amines** are classified as **primary, secondary, or tertiary** according as one, two or three hydrogen atoms in the ammonia molecule have been substituted by alkyl groups.

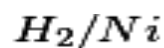


**3. Preparation of amines:**

(i) By reduction of nitro compounds.



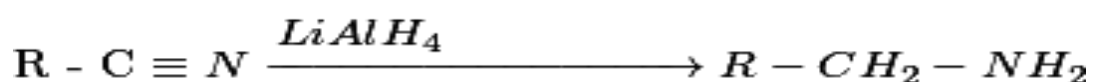
(ii) By reduction of nitriles.



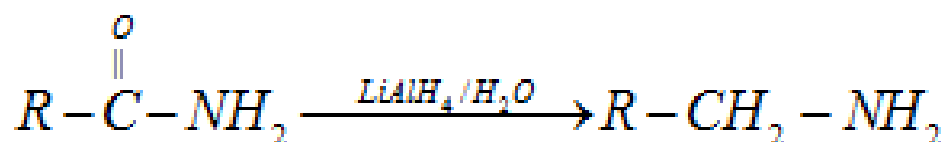
Or



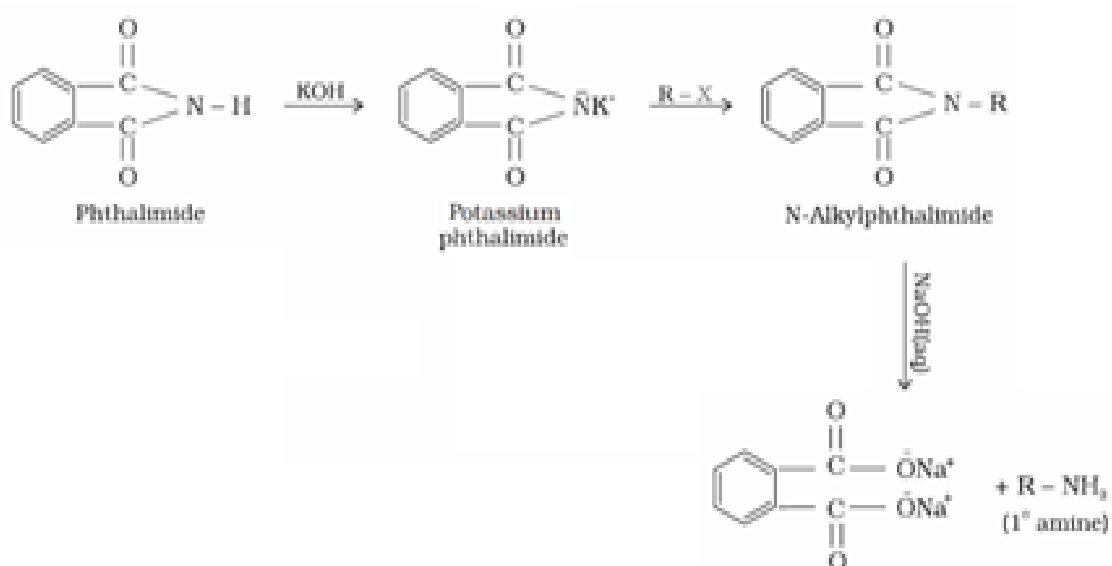
Or



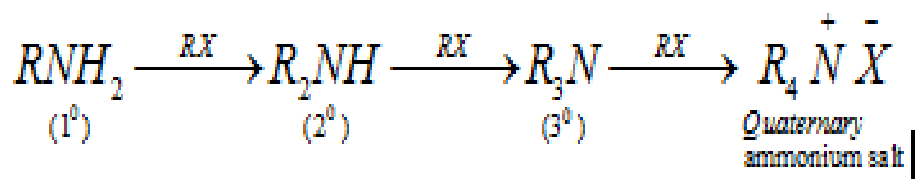
(iii) By reduction of amides.



(iv) By Gabriel phthalimide synthesis: For primary alkyl amines only, not for aromatic amines.



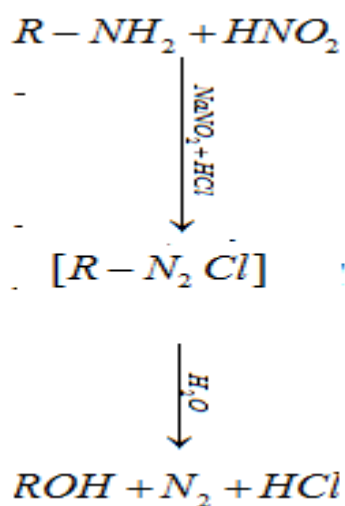
(v) By Hoffmann bromide degradation reaction



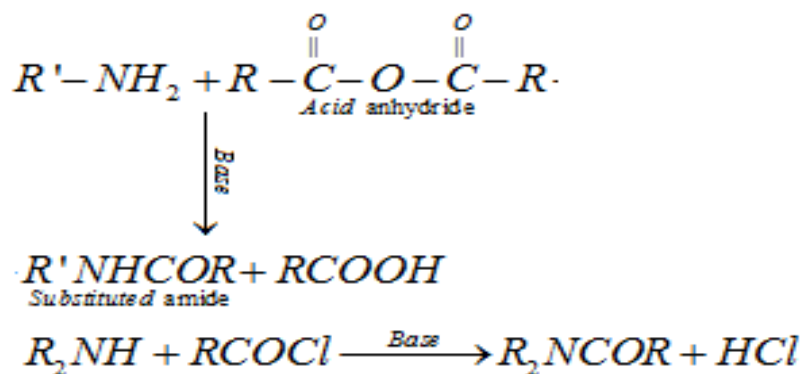
4. All the three classes of aliphatic amines (1°, 2° and 3°) form H-bonds with water. As a result, lower aliphatic amines are soluble in water.

5. Some important reactions of amines

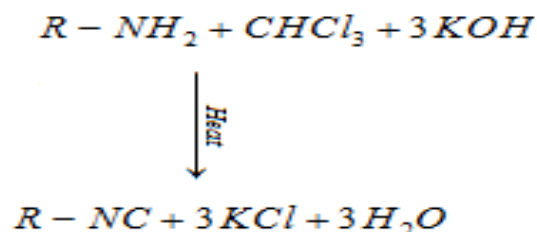
(i) Amines are basic in nature they react with acid to form salt.



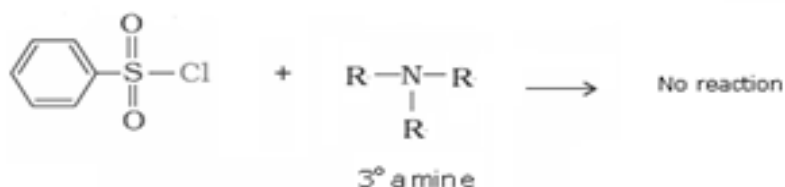
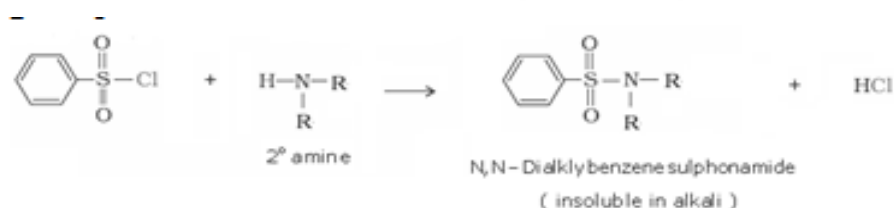
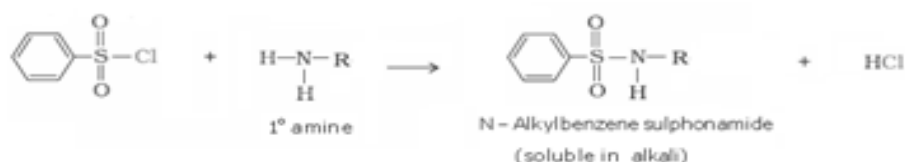
(ii) Aliphatic and aromatic primary and secondary amines react with acid chlorides, anhydrides and esters by nucleophilic substitution reaction. This reaction is known as **acylation**.



(iii) Aliphatic and aromatic primary amines on heating with chloroform and ethanolic potassium hydroxide form isocyanides or carbylamines which are foul smelling substances. Secondary and tertiary amines do not show this reaction. This reaction is known as **carbylamine reaction** or **isocyanide test** and is used as a test for primary amines.



(iv)  $\text{C}_6\text{H}_5\text{SO}_2\text{Cl}$  is also called **Hinsberg's reagent** and reacts with primary and secondary amines to form sulphonamides. Tertiary amines do not react with benzene sulphonyl chloride.



**6. Diazonium salt** have the general formula  $\text{RN}_2^+ \text{X}^-$ , where R stands for an aryl group and  $\text{X}^-$  ion maybe  $\text{Cl}^-$ ,  $\text{Br}^-$ ,  $\text{HSO}_2^-$ ,  $\text{BF}_4^-$ , etc.

**7. Primary aliphatic amines** form highly unstable alkyl diazonium salts which decomposes to give alcohols with the evolution of  $\text{N}_2$ . Primary aromatic amines form arene diazonium salts which are stable for a short time in solution at low temperature (273 -278 K).

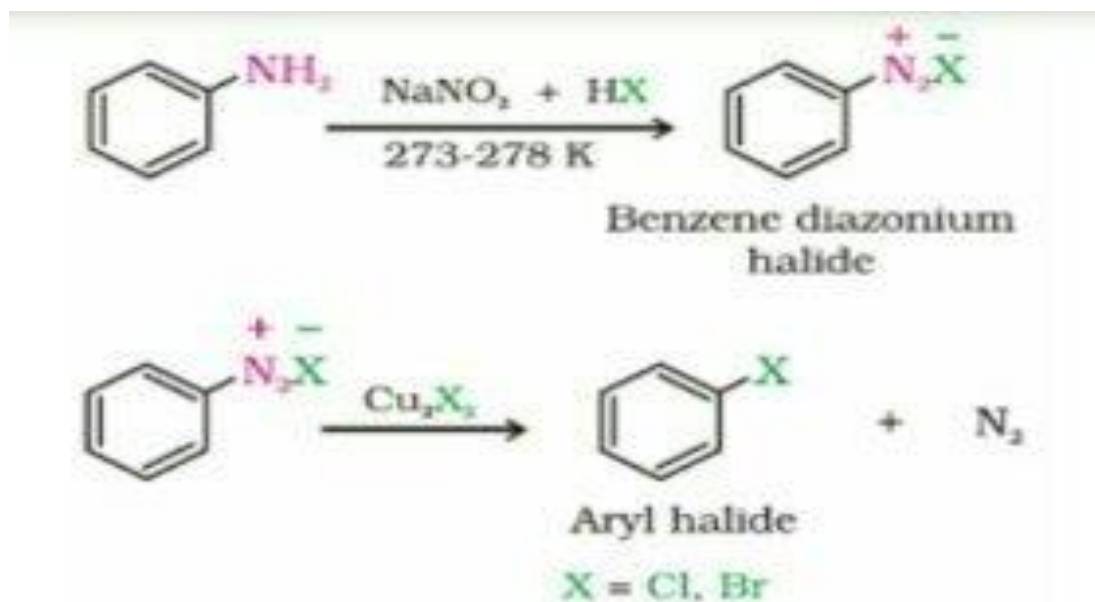
**8. Arenediazonium ion** is resonance stabilized.

9. Benzene diazonium chloride is prepared by the reaction of aniline with nitrous acid at 273 – 278 K. This conversion of primary aromatic amines into diazonium salt is known as **diazotisation**.

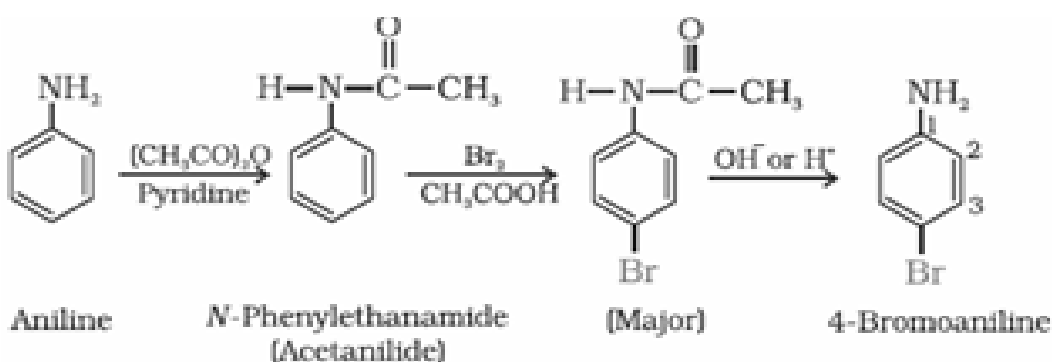
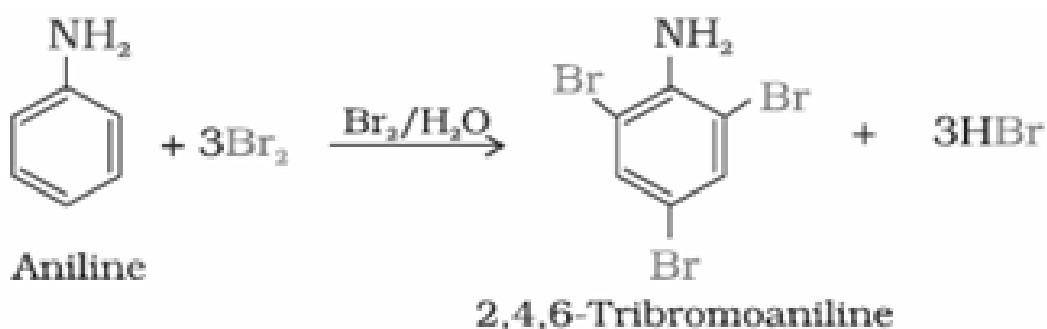
### 10. Chemical properties:

(i) Diazonium group being a very good leaving group, is substituted by other groups such as  $\text{Cl}^-$ ,  $\text{Br}^-$ ,  $\text{I}^-$ ,  $\text{CN}^-$  and  $\text{OH}^-$ .

(ii) The  $\text{Cl}^-$ ,  $\text{Br}^-$  and  $\text{CN}^-$  nucleophiles can easily be introduced in the benzene ring in the presence of Cu (I) ion. This reaction is called **Sandmeyer reaction**.



(iii) **Bromination:** Aniline reacts with bromine water at room temperature to give a white precipitate of 2, 4, 6-tribromoaniline



In order to stop reaction at mono-substitution activating effect of NH<sub>2</sub> group is reduced by acetylation. This prevents di and tri substituted products. Acetyl group is removed by hydrolysis.

**11. Importance of diazonium salts:** They are very good intermediates for the introduction of -F, -Cl, -Br, -I, -CN, -OH, -NO<sub>2</sub> groups into the aromatic ring.

