CONCEPTS -

* **Adsorption** :- The accumulation of molecular species at the surface rather than in the bulk of a solid or liquid is termed as **adsorption**.

* **Adsorbate** :- The molecular species which concentrates or accumulates at the surface is termed as **adsorbate**.

* **Adsorbent** :- The material on ~~which~~ the surface of which the adsorption takes place is called **adsorbent**.

* **Desorption** :- The process of removing an absorbed substance from a surface on which it is adsorbed is called **desorption**.

* **Absorption** :- It is the phenomenon in which a substance is uniformly distributed all over the surface.

* **Sorption** :- When adsorption and absorption take place simultaneously.

Catalysis :- Substances which accelerate the rate of ~~reaction~~ chemical reaction, and themselves remain chemically and quantitatively unchanged after the reaction are known as catalysts and phenomenon is known as catalysis.

i) Homogenous catalyst :- when reactants, products and catalyst are in same phase.

Ex, $2SO_2(g) + O_2(g) \xrightarrow{NO(g)} 2SO_3(g)$.
(Lead chamber process)

ii) Heterogenous catalyst :- when reactants, ~~products~~ and catalysts are in different phases is known as heterogeneous catalyst.

Ex, $N_2(g) + 3H_2(g) \xrightarrow{Fe(s)} 2NH_3(g)$.
(Haber's process)

* Colloids :- A colloid is a heterogeneous system in which one substance is dispersed as very fine particles in another substance, called dispersion medium.