

## FORMULA

\* Formal charge on an atom in a Lewis structure

$$= V - N - \frac{1}{2} B$$

$\therefore V =$  Total no. of valence electrons in free atom.

$N =$  Total no. of non bonding (lone pair) electrons.

$B =$  Total no. of bonding ~~at~~ (shared) electrons.

### Rules

\* To select the best structure from formal charge.

i) The structure with formal charge of zero-preferred.

ii) The -ve charge should be on atom of more electronegative element.