

* * Chemical Bond — is a bond which holds atoms or ions ^{together} in the molecule by a strong combining force.

* There are two types of chemical bonds :-

• Ionic Bond :- Ionic bond is the bond formed by a strong force of attraction between two oppositely charged ions.

Ex, NaCl, \bar{I}

• Covalent Bond :- Bond formed by sharing of electrons between two combining atoms in outermost orbit to achieve octet configuration.

Ex, O₂, N₂, H₂.

* Electrovalence = Number of unit charges on the ion.

Following steps to represent Lewis Dot structures.

Step 1. Determine the no. of valence electrons.

Step 2. Determine the central atom

Step 3: Then subtracts the no. of electrons consumed for forming single bonds from the total valence electrons.

Step 4:- Distribute the remaining ~~no~~ valence electrons as pairs around each terminal atoms so for each atom attains octet rule.

Step 5:- Match the no. of electrons which was distributed to the no. of valence electron counted in Step 1.

Step 6:- Complete the octet on the central atom.