
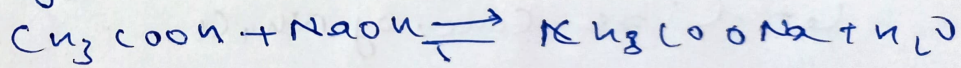


Q The pH of a solution obtained by mixing 100 mL of 0.2 M CH_3COOH with 0.1 M NaOH would be:

Soln 

Conc. of salt formed = 0.1 M



$$\text{pH} = \frac{1}{2} [\text{p}K_w + \text{p}K_a + \log C]$$

$$= \frac{1}{2} [14 + 4.74 -] = \boxed{8.87}$$