

The aqueous solution of sugar does not conduct electricity. However, when sodium chloride is added to water, it conducts electricity. How will you explain this statement on the basis of ionisation and how is it affected by concentration of sodium chloride?

- (i) Sugar being a non-electrolyte does not ionize in water, whereas NaCl ionizes completely in water and produces Na^+ and Cl^- ions which help in the conduction of electricity.
- (ii) When concentration of NaCl is increased, more Na^+ and Cl^- ions will be produced. Hence, conductance increases