

K_a for CH_3COOH is 1.8×10^{-5} and K_b for NH_4OH is 1.8×10^{-5} . The pH of ammonium acetate will be

- (i) 7.005
- (ii) 4.75
- (iii) 7.0
- (iv) Between 6 and 7

Sol: (c) Ammonium acetate is a salt of weak acid and weak base.