

**Assertion (A) :** Increasing order of acidity of hydrogen halides is  
 $\text{HF} < \text{HCl} < \text{HBr} < \text{HI}$

**Reason (R) :** While comparing acids formed by the elements belonging to the same group of periodic table, H-A bond strength is a more important factor in determining acidity of an acid than the polar nature of the bond.

- (i) Both A and R are true and R is the correct explanation of A.
- (ii) Both A and R are true but R is not the correct explanation of A.
- (iii) A is true but R is false.
- (iv) Both A and R are false.

Ans : i