

Diborane ( $B_2H_6$ ) reacts independently with  $O_2$  and  $H_2O$  to produce, respectively.

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- (a)  $B_2O_3$  and  $H_3BO_3$
- (b)  $B_2O_3$  and  $[BH_4]^-$
- (c)  $H_3BO_3$  and  $B_2O_3$
- (d)  $HBO_2$  and  $H_3BO_3$

It gets hydrolysed readily to give boric acid.

