

Ionisation enthalpy ( $\Delta_i H_1$ ,  $\text{kJ mol}^{-1}$ ) for the elements of Group 13 follows the order.

- (i)  $B > Al > Ga > In > Tl$
- (ii)  $B < Al < Ga < In < Tl$
- (iii)  $B < Al > Ga < In > Tl$
- (iv)  $B > Al < Ga > In < Tl$

Ans : iv

Reason : Boron have small size and large charge density as a result have high IE and Gallium shows inert pair effect

Learn as exception