

20. Which one of the following pairs of molecules will have permanent dipole moments for both members? [2003]
- (a)  $\text{NO}_2$  and  $\text{CO}_2$                       (b)  $\text{NO}_2$  and  $\text{O}_3$   
(c)  $\text{SiF}_4$  and  $\text{CO}_2$                       (d)  $\text{SiF}_4$  and  $\text{NO}_2$
- 

20. (b) Both  $\text{NO}_2$  and  $\text{O}_3$  have angular shape and hence will have net dipole moment.
-