

**Q.14** Given below are two statements : one is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : Dipole-dipole interactions are the only non-covalent interactions, resulting in hydrogen bond formation.

Reason R : Fluorine is the most electronegative element and hydrogen bonds in HF are symmetrical.

In the light of the above statements, choose the most appropriate answer from the options given below :

- A Both A and R are true and R is the correct explanation of A

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- B A is true but R is false

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- C A is false but R is true

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- D Both A and R are true but R is NOT the correct explanation of A

**26th Feb Morning Shift 2021**

**Ans 14.** Dipole - Dipole are not only the interaction responsible for hydrogen bond formation. Ion-dipole can also be responsible for hydrogen bond formation. F is most electronegative element and anhydrous HF in solid phase has symmetrical hydrogen bonding.