

➔ For an endothermic reaction, where ΔH represents the enthalpy of the reaction in kJ/mol , the minimum value for the energy of activation will be [IIT 1992]

- A) Less than ΔH
- B) Zero
- C) More than ΔH
- D) Equal to ΔH

Correct Answer: C

Solution :

For a reaction E_a for forward reaction = E_a for backward reaction + ΔH .