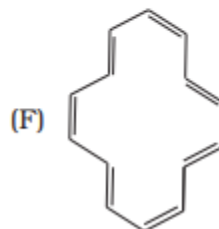
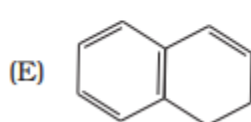
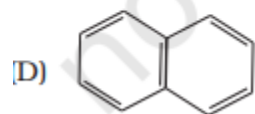
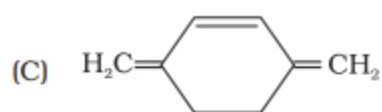
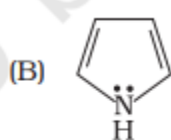
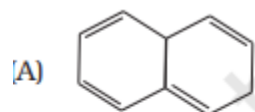


Which of the following compounds are aromatic according to Huckel's rule?



- A = Has  $8\pi$  electrons, does not follow Huckel rule. Orbitals of one carbon atom are not in conjugation. It is not aromatic.
- B = Has  $6\pi$  delocalised electrons. Hence, is aromatic.
- C = Has  $6\pi$  electrons in conjugation but not in the ring. Non aromatic.
- D =  $10\pi$  electrons in planar rings, aromatic.
- E = Out of  $8\pi$  electrons it has delocalised  $6\pi$  electrons in one six membered planar ring, which follows Huckel's rule due to which it will be aromatic.
- F =  $14\pi$  electrons are in conjugation and are present in a ring. Huckel's rule is being followed. Compound will be aromatic if ring is planar.