

- The reaction $2\text{NO}_2(\text{g}) \rightleftharpoons \text{N}_2\text{O}_4(\text{g})$ is an exothermic equilibrium. This means that :
- (a) equilibration of this gas mixture will be slower at high temperature
 - (b) a mole of N_2O_4 will occupy twice the volume of a mole of NO_2 at the same?
 - ~~(c)~~ the equilibrium will move to the right if an equilibrium mixture is cooled
 - (d) the position of equilibrium will move to the left with increasing gas pressure

(c) Temperature \downarrow favours exothermic reaction.