Consider the following reactions. In which case the formation of product is favoured by decrease in pressure?

(1)
$$CO_2(g) + C(s) \rightleftharpoons 2CO(g)$$
;

$$\Delta H^{\circ} = +172.5 \text{ kJ}$$

(2)
$$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$$
;

$$\Delta H^{\circ} = -91.8 \text{ kJ}$$

(3)
$$N_2(g) + O_2(g) \rightleftharpoons 2NO(g)$$
;

$$\Delta H^{\circ} = 181 \,\mathrm{kJ}$$

(4)
$$2H_2O(g) \iff 2H_2(g) + O_2(g)$$
;

$$\Delta H^{\circ} = 484.6 \text{ kJ}$$

1, 4

(d) Pr. $\downarrow \Rightarrow$ towards more gaseous moles