

86. Given the following reaction at equilibrium, $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$. At constant pressure is some inert gas added in to the system. Predict which of the following facts will be affected?

- (a) More $\text{NH}_3(\text{g})$ is produced
- (c) No affect on the equilibrium

- (b) Less $\text{NH}_3(\text{g})$ is produced
- (d) K_p of the reaction is decreased

(b) Addition of inert gas at constant pressure will increase volume so reaction moves in the direction having more gaseous moles.

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