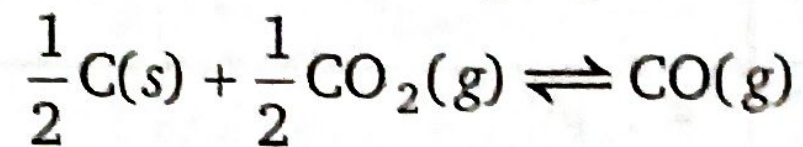
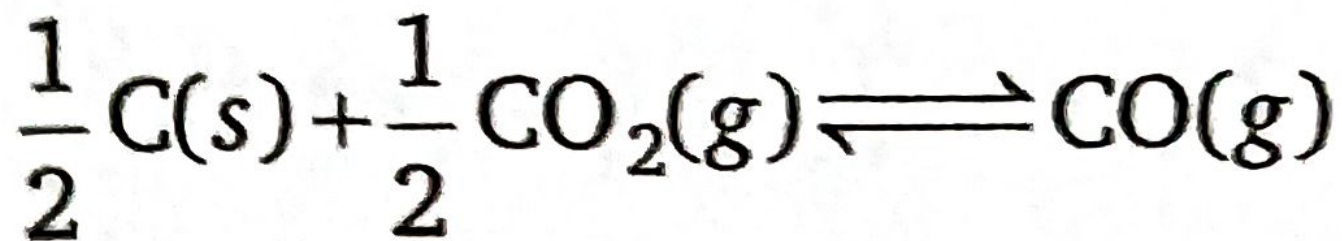


1. If 50% of CO_2 converts to CO at the following equilibrium :



and the equilibrium pressure is 12 atm. Calculate K_p .

4



$t = 0$

$P \text{ atm}$

$t_{\text{eq.}}$

$(P - P/2)$

P

$$P/2 + P = 12$$

$$P = 8 \text{ atm}$$

$$K_p = \frac{(p_{\text{CO}})^1}{(p_{\text{CO}_2})^{1/2}}$$