

Question 21: What is the oxidation number of Br in KBrO_4

ANSWER: OPTION 4

Total charge=0.

K is a group I element. So, its oxidation no. =+1.

Oxidation no. of oxygen =-2.

Let, Oxidation no. of Br be x.

So, $1+x+4(-2)=0$

$$1+x-8=0$$

$$x-7=0$$

$$\mathbf{x= +7}$$

So, oxidation state of Bromine is +7.