

Question 19: Molecular weight of oxalic acid is 126. The weight of oxalic acid required to neutralise 100 cc of 1 normal solution of NaOH is

ANSWER: OPTION 1

no. of equivalent of acid = no of equivalent of Base

$$[(\text{Weight}/\text{Molecular weight}) \times \text{n-factor}]_{\text{Base}} = (\text{N} \times \text{Vlitre})_{\text{Acid}}$$

∴ n factor of NaOH=1

∴ n factor of Oxalic acid =1

Let W be the weight of oxalic acid required

Therefore,

$$= [W \times 1000 / 63 \times 100]. W = 1$$

$$= \mathbf{W=6.3 \text{ gm}}$$