

Question 18: Volume of 0.1 M H₂SO₄ required to neutralize 30 mL of 0.2 N NaOH is

ANSWER: OPTION 1

0.1M of H₂SO₄ ⇒ 0.2N of H₂SO₄

$$\therefore N_1V_1 = N_2V_2$$

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$$0.2 \times V_1 = 0.2 \times 30$$

$$\mathbf{V_1 = 30ml}$$