Question 14: 20 ml of 0.2 M NaOH (aq) solution is mixed with 35 ml of 0.1 M NaOH (aq) solution and the resultant solution is diluted to 100 ml. 40 ml of this diluted solution reacted with 10% impure sample of oxalic acid H2C2O4. The weight of the impure sample is

ANSWER: OPTION 1

M_{NaOH} (resultant) = (20×0.2+35×0.1)/100

$\qquad$ = 0.075M

Milli-equivalent of NaOH = Milli-equivalent of $H_2C_2O_4$

Let weight of impure sample is X g.

$\qquad$ 40 × 0.075 = (X × 0.90)/100 × 2 × 1000

 ANSWER= $\qquad$ **X=0.15g**