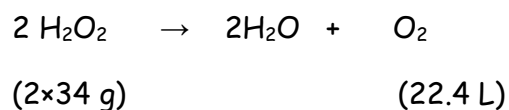## Question 2:

The volume strength of 1.5 N H2O2 is

Answer (option 2)   8.4

 Equivalent weight of H_2O_2 is 17

Normality (N) = 1.5

Strength of H_2O_2 = Normality x Equivalent weight = 1.5 ×17 = 25.5

$2 H_2O_2 \quad \rightarrow \quad 2H_2O \; + \quad O_2$

(2×34 g)   (22.4 L)

Since 68 grams of H_2O_2 produces 22.4 litres of oxygen at NTP,

therefore, 25.5 grams of H_2O_2 will produce

= 22.4/68 x 25.5 = 8.4 litre of oxygen

Thus, volume strength of given H_2O_2 solution is 8.4.