

## Previous Year JEE Problems

3) Amongst LiCl, RbCl, BeCl<sub>2</sub> and MgCl<sub>2</sub> the compounds with the greatest and the least ionic character, respectively are :

- (1) RbCl and MgCl<sub>2</sub>
- (2) LiCl and RbCl
- (3) MgCl<sub>2</sub> and BeCl<sub>2</sub>
- (4) RbCl and BeCl<sub>2</sub>

Solution: Because of the smallest cationic size BeCl<sub>2</sub> is least ionic. Rb<sup>+</sup> has the biggest ionic size. So it has the greatest ionic character. Hence option (4) is the answer.

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