

Points to Remember

• Entropy : Pure Solvent $<$ Solution

• Boiling point of pure solvent is always less than Boiling point of solution. That's why, we mention ELEVATION in B.P. (ΔT_b) of solution after adding solute.

• Freezing point of pure solvent is always greater than Freezing point of solution. That's why, we mention DEPRESSION in FP (ΔT_f) of solution after adding solute.

→ Note: Therefore, it is difficult to boil a solution or to freeze a solution with respect to pure solvent