Points to Remember

- K_H (Henry constant) depends upon nature of solute, nature of solvent and temperature.
  Also, on increasing temperature, $K_H \uparrow$, solubility $\downarrow$
- Vapour pressure of a pure liquid depends only on temperature and nature of liquid.
- Note: Do practice questions based on application of Raoult law and Dalton's law as these topics are very important from exam point of view.

$$P_{Total} = P_A^\circ X_A + P_B^\circ X_B$$

where P_A° = VP of pure liquid A
P_B° = VP of pure liquid B.

$$P_A = P_T Y_A = P_A^\circ X_A$$

$P_T Y_A$ ⇓ from Dalton's law

$P_A^\circ X_A$ ⇓ from Raoult's law.

$$\frac{1}{P_T} = \frac{Y_A}{P_A^\circ} + \frac{Y_B}{P_B^\circ}$$

And Also, $X_A + X_B = 1$

$Y_A + Y_B = 1$