Points To Note:-

* Assumptions of KTG:-

→ All gases are made of molecules moving in random directions.

→ Size of molecules is much smaller than average separation between them.

→ Molecules exert no force on each other except during collison.

→ All collisons between molecules and between molecules and wall is elastic.

→ Molecules obey Newton's laws of motion.

→ Density and distribution of molecules with different velocities is independent of space, time and direction.

→ For these assumptions, to be true, no. of molecules should be very-very large, pressure should be low and temperature should be high.

Ideal gases:- Gases which obey ideal gas laws.

Real gases:- They deviate atleast slightly from ideal gas laws because of two factors:-

→ Gas molecules attract one another.

→ Gas molecules occupy a finite volume.

* Ideal gases exist at high temperature and low pressure.