

TIPS:

- Read NCERT at least twice and make short notes.
- Revise Notes Regularly.
- Attempt as many PYQ's as possible.
- Structures of basic compounds should be on your tips.

TRICKS:

An easy way to calculate the Bond order (up to atomic number 18)

Drawing the molecular orbital diagram every time in order to calculate the bond order can be time consuming. But a small trick can help us solve "bond order type" questions in no time. In this, you should know the atomic number of the atoms in the molecules and you are all set. Here you go:

Add the **atomic numbers** of the atoms in the molecule.

Match the value with the bond order in the table below. If your value comes out to be **14**, this means the bond order is **3**.

You have to be a lot more careful if it's a charged species. Say, we are talking about a charged species with a charge of **+2**. If the sum of the atomic numbers comes out to be 14, then subtract **2** from **14**. The new value come out to be 12, so the bond order is 2.

In case of a **negative charged** species, we do the opposite. We **add** the charge to the value of the sum of the atomic numbers and match the new value with the bond order in the table.

Always remember this trick is useful in the case when the total atomic number of the molecule is between 10 and 18.

Total Atomic Number	Bond Order
10	1
11	1.5
12	2
13	2.5
14	3
15	2.5
16	2
17	1.5
18	1