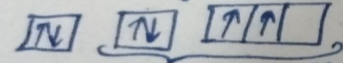


* Hybridisation: Mixing of atomic orbitals and redistribution of orbitals in specific orientation.

1.) Necessary Condition: orbitals in valence shell, low energy difference; promotion of e^- is not necessary, both half filled & fully filled orbitals can undergo hybridisation.

2.) Outcomes: (i) No. of orbitals that are hybridized = No. of hybridized orbitals obtained.
(ii) Hybrid orbitals formed are degenerate & orient in specific directions.

3.) $C = 1s^2 2s^2 2p^2$ Tetra valency



sp, sp^2, sp^3 Hybridization

