

Which of the following oxidation States to group 14 elements exhibit?

- a) +1, +5
- b) +5, +2
- c) +2, +4
- d) +3, +5

Answer: c

Explanation: The elements of carbon family exhibit + 2 and + 4 oxidation state. The compounds of plumbum in + 4 oxidation state are powerful oxidizing agents since +2 oxidation state of plumbum is more stable due to inert pair effect.