

For the reaction, $2\text{N}_2\text{O}_5 \rightarrow 4\text{NO}_2 + \text{O}_2$, the rate equation can be expressed in two ways $-\frac{d[\text{N}_2\text{O}_5]}{dt} = k[\text{N}_2\text{O}_5]$

and $+\frac{d[\text{NO}_2]}{dt} = k'[\text{N}_2\text{O}_5]$

k and k' are related as: **[Online April 11, 2014]**

- (a) $k = k'$ (b) $2k = k'$
(c) $k = 2k'$ (d) $k = 4k'$