

The differential rate law for the reaction
 $\text{H}_2 + \text{I}_2 \rightarrow 2\text{HI}$ is

[2002]

(a) $-\frac{d[\text{H}_2]}{dt} = -\frac{d[\text{I}_2]}{dt} = -\frac{d[\text{HI}]}{dt}$

(b) $\frac{d[\text{H}_2]}{dt} = \frac{d[\text{I}_2]}{dt} = \frac{1}{2} \frac{d[\text{HI}]}{dt}$

(c) $\frac{1}{2} \frac{d[\text{H}_2]}{dt} = \frac{1}{2} \frac{d[\text{I}_2]}{dt} = -\frac{d[\text{HI}]}{dt}$

(d) $-2 \frac{d[\text{H}_2]}{dt} = -2 \frac{d[\text{I}_2]}{dt} = \frac{d[\text{HI}]}{dt}$