◆**What will be the effect of temperature on rate constant?**

Ans:

A 10^0 increase in temperature almost doubles the rate constant of a process. The Arrhenius equation, on the other hand, gives the exact temperature dependency of the rate of a chemical reaction.

$k = Ae^{-Ea/RT}$

Where A stands for the Arrhenius factor, also known as the frequency factor.
T stands for temperature.
The gas constant is R.

The activation energy is referred to as E_a.