

24. Rate of a reaction can be expressed by Arrhenius equation as: $k = Ae^{-E/RT}$ In this equation, E represents (2006)

- 1) the energy above which all the colliding molecules will react
- 2) the energy below which colliding molecules will not react
- 3) the total energy of the reacting molecules at a temperature, T
- 4) the fraction of molecules with energy greater than the activation energy of the reaction

Ans. (2) Colliding molecules with energy less than activation energy will not react.